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$$\Delta H^\circ = -92 \text{ kJ mol}^{-1}$$
$$\Delta S^\circ = 192 \text{ J mol}^{-1} \text{ K}^{-1}$$
To find the Gibbs free energy of formation at room temperature, recall that $G^\circ = H^\circ - T \Delta S^\circ$
$$G^\circ = -92 \text{ kJ mol}^{-1} + (298 \text{ K})(-0.192 \text{ kJ mol}^{-1} \text{ K}^{-1}) = -35 \text{ kJ mol}^{-1}$$
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