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CHEM113L:
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Determining an
Equilibrium
Constant by
Spectrophotometry
Procedure Le
Chatelier's
Principle Lab with
Cobalt Complex
Ions Determination
of K_{eq} for
 $FeSCN_2^+$ Lab
Explanation Video

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~~ANSWERS: Cobalt~~

~~CoCl₄²⁻ and~~

~~Cu(H₂O)₆²⁺ Lab~~

Experiment #13:

The Equilibrium

Constant.

FeSCN₂⁺ Equilbriu

m - LeChatelier's

Principle Lab Part 1

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ANSWERS: Fe³⁺

and FeSCN₂⁺

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Principle of
Chemical

Equilibrium - Basic
Introduction

Equilibrium: Crash
Course Chemistry

#28 Equilibrium

Lab ANSWERS: Chr
omate/Dichromate

Keq FeSCN₂+ Lab
Equilibrium

Constant

The Effect of

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Pressure on
Equilibrium - N_2O_4
to $2NO_2$

Amazing

*Experiments with
Water Equilibrium*

*Equations: Crash
Course Chemistry*

*#29 Equilibrium in
cobalt(II) chloride*

solution VCE UNIT

4: Equilibrium Iron

Thiocyanate Blue

Bottle Equilibrium

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Effect of

Temperature on an
Equilibrium

Reaction The
Equilibrium

Constant Chemical
equilibrium of

liquid nitrogen
dioxide Iron(III)

Nitrate and
Potassium

Thiocyanate

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Lab: Determination

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*of an Equilibrium
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~~10. Chemical
Equilibrium~~

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CONSTANT Pre-Lab
- NYB Chemistry of
Solutions

~~Equilibrium~~

~~Constant Lab Part~~

~~1: K, Beer's Law,
and Stoichiometry~~

~~Chemical~~

~~equilibrium with~~

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~~real examples~~

Chromate Dichromate Ion Equilibrium - LeChatelier's Principle Lab Part 2

~~NO₂ N₂O₄ Gas
Equilibrium-~~

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Effect of

Temperature on
Cobalt Complex

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Balance Lab

Answers ...

Restoring Balance

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tube B, which was

a combination of

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distilled water and cobalt chloride, was an endothermic reaction. It was an endothermic reaction because when the test tube was heated, the solutions turned from pink to dark blue.

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and the FeSCN_2 ...

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markets chemicals
in a manner that
delivers excellence
and value to all of
our clients. The
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improvement and
to exploring new
ideas to serve our
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Châtlier's Principle

It has been

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observed that
when a reaction at
equilibrium is
disturbed by
changing either the
concentration of
one of the chemical
components, the
total pressure, or
the temperature,
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about 100 mL of water on a hot plate. While waiting for the water to boil, move on to the next step.

3. Put on latex gloves and place five drops of CoCl_2 solution in each of the 24 wells of the labeled well plate.
- 4.

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escape, shifting the
equilibrium, below,
to the right, and
the soda goes flat:
 $\text{CO}_2(\text{aq}) \rightleftharpoons \text{CO}_2(\text{g})$
Another possibility
is to chemically
react away one of

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the products. If you remove products from the reaction system, the equilibrium will shift to the right. The reverse is also true; if reactants are removed, the equilibrium will shift to the left. 3.

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/ base titration

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